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The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N_2 are mixed with 12.0 mol of H_2 according to the following equation: $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$...

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Stoichiometry SECTION 2 PROBLEMS Write the answer on the line to the left Show all your work in the space provided 1 45 mol The following equation represents a laboratory preparation for oxygen gas: $2KClO_3(s) \rightarrow 2KCl(s) + 3O_2(g)$ How many moles of O_2 form if 30 mol of

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Calculate the molar mass of the following, and then write the molar mass in ratio form. KNO_3 . Na_2SO_4 . $Ca(OH)_2$ $(NH_4)_2SO_3$. $CuSO_4$. Moles and Mass. Determine the ...

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Chapter 9: Standard Review Worksheet

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